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Chemistry

- A cathode ray tube is made of glass containing two thin pieces of metal called-
(A) Cathode
(B) Anode
(C) Electrode
(D) None of these
- In the presence of an electric or magnetic field, how do cathode rays behave?
(A) Positive
(B) Negative
(C) Neutral
(D) Both (A) and (B)
- The ratio of Charge and Mass of an electron is-
(A) $1.758820 \times 10^{11} \text{ C kg}^{-1}$
(B) $1.058820 \times 10^{13} \text{ C kg}^{-1}$
(C) $2.758820 \times 10^{14} \text{ C kg}^{-1}$
(D) $1.838820 \times 10^{10} \text{ C kg}^{-1}$
- In chemistry, atomic theory is a scientific theory of the nature of matter, which states that matter is composed of discrete units called _____.
(A) Atoms
(B) Electrodes
(C) Nucleus
(D) Quarks
- Which of the following objects cannot be changed by changing its temperature and pressure?
(A) wood
(B) water
(C) ideal gas
(D) None of these
- When a substance is converted directly from solid to gas on heating, this process called –
(A) Sublimation (B) Distillation
(C) Evaporation (D) Condensation
- Which of the following is the fourth fundamental state of matter, and was first described by chemist Irving Langmuir in the 1920s?
(A) Plasma
(B) Bose-Einstein Condensate
(C) LHC
(D) None of these
- The inorganic solid mixture is separated by which of the following methods?
(A) Distillation Method
(B) Chromatography
(C) Sublimation
(D) Crystallisation
- When the difference in boiling points of two liquids is higher, then by which method their mixture is separated?
(A) Fractional Distillation
(B) Steam Distillation
(C) Distillation
(D) Chromatography
- The temperature at which a matter converts to its liquid state from its solid state, is called-
(A) Melting point
(B) Boiling point
(C) Freezing point
(D) None of these
- What is the effect of a substance on both freezing point and melting point in the presence of impurities?
(A) both freezing point and melting point increases
(B) both freezing point and melting point decreases
(C) freezing point increases and melting point decreases
(D) freezing point decreases and melting point increases

12. The action of evaporation depends on which of the following?
 (A) Fluid temperature
 (B) Open surface area of fluid
 (C) On the surface of the fluid
 (D) All of the above
13. Which of the following mixtures can be separated by sublimation process?
 (A) Naphthalene and anthracene
 (B) Camphor and Zinc
 (C) Ammonium Chloride and Iodine
 (D) All of the above
14. What is the effect of increasing the pressure on the fluid state of a substance on its boiling point?
 (A) increases
 (B) decreases
 (C) no effect
 (D) None of these
15. Chemical formula of bauxite is?
 (A) $\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O}$
 (B) $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$
 (C) $\text{Al}_2\text{O}_3 \cdot 3\text{H}_2\text{O}$
 (D) $\text{Al}_2\text{O}_3 \cdot 4\text{H}_2\text{O}$
16. Which of the following quantum numbers represents sub energy levels?
 (A) Prime quantum number
 (B) Dual quantum numbers
 (C) Magnetic quantum number
 (D) Rotation quantum number
17. Which one of the following sentences is the correct definition of Pauli's exclusion rule -
 (A) For any two electrons in an atom, the value of the four quantum numbers cannot be the same.
 (B) Electrons are not coupled until a blank orbital is attainable
 (C) Simultaneous exact determination of position and velocity of a particle is impossible
 (D) If the value of the four quantum numbers is the same for an electron, then its rotation will be the same with respect to the other electron.
18. Which of the following is the most isotopic element?
 (A) Hydrogen
 (B) Rubidium
 (C) Cesium
 (D) Polonium
19. The formula for ultimate temperature is –
 (A) $T = -273^\circ + t^\circ\text{C}$ (B) $T = 273^\circ + t^\circ\text{C}$
 (C) $T = 273^\circ + t^\circ\text{F}$ (D) $T = -273^\circ + t^\circ\text{F}$
20. Which of the following is not an organic compound?
 (A) C^3H^8 (B) $\text{C}^6\text{H}^5\text{OH}$
 (C) $\text{CH}_3(\text{CH}_2)_7\text{CH}_3$ (D) H_2CO_3
21. Which of the following concept of acids is given by Bronsted Lowery?
 (A) An acid is a molecule or ion which is capable of donating a proton.
 (B) An acid is a substance which can accept an electron
 (C) Acid is a substance which produces hydrogen ions in aqueous solution
 (D) Acid is a substance which can produce an electron
22. Some bases are insoluble or partly soluble in water such bases are called weak bases. Which of the following is a weak base?
 (A) potassium hydroxide
 (B) sodium hydroxide
 (C) ammonium hydroxide
 (D) Barium hydroxide
23. Which acid is found in tomato?
 (A) Tartaric acid (B) Oxalic acid
 (C) Malic Acid (D) Lactic acid
24. Which of the following alkali is found in window cleaners?
 (A) Ammonium hydroxide
 (B) Potassium hydroxide
 (C) Magnesium hydroxide
 (D) Calcium hydroxide
25. Which of the following is correct for Oxidation?
 (A) The loss of electrons or an increase in oxidation state by a molecule, atom, or ion.
 (B) The gain of electrons or a decrease in oxidation state by a molecule, atom, or ion.
 (C) The loss of electrons or a decrease in oxidation state by a molecule, atom, or ion.
 (D) The gain of electrons or an increase in oxidation state by a molecule, atom, or ion.
26. Which of the following is not a matter?
 (A) Water (B) Ice
 (C) Light (D) Oxygen
27. What is boiling point of nitrogen?
 (A) -189.6°C (B) -200.4°C
 (C) -195.8°C (D) -205.3°C

28. How many molecules are there in the volume of one mole of any substance?
 (A) 6.021×10^{25} (B) 6.022×10^{23}
 (C) 6.243×10^{23} (D) 6.258×10^{24}
29. How much number of Atoms present in 12g of Carbon of C-12 Isotope?
 (A) 6.026×10^{25} (B) 6.022×10^{23}
 (C) 6.943×10^{23} (D) 6.457×10^{24}
30. Which is the most malleable element?
 (A) Platinum (B) Silver
 (C) Gold (D) Copper
31. Which of the following is highest electro-negative element?
 (A) Oxygen (B) Neon
 (C) Sodium (D) Fluorine
32. Who is the first person to discover the evidence of Radio- Activity?
 (A) Antoine Henri Becquerel
 (B) Marie Curie
 (C) Alfred Nobel
 (D) Dmitri Ivanovic Mendeleev
33. What is the density of pure water at 4 degree Celsius?
 (A) 1.3 gm/cm^3 (B) 1.1 gm/cm^3
 (C) 1 gm/cm^3 (D) 1.05 gm/cm^3
34. What is the effect on surface tension if we increase the temperature?
 (A) Decrease
 (B) Increase
 (C) No Effect
 (D) Depend on Atmosphere
35. What is atomic number?
 (A) Number of electrons
 (B) Number of protons
 (C) Number of neutrons
 (D) None of these
36. The chemical reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ $\Delta H^\circ = -91.8 \text{ kJ/mol}$ is called-
 (A) Wurtz reaction
 (B) Ostwald Process
 (C) Habar Process
 (D) None of these
37. What is the mass of Oxygen?
 (A) 18.241 u (B) 15.998 u
 (C) 16.999 u (D) 14.958 u
38. What is molar mass of NaOH?
 (A) 36.997 g/mol (B) 40.997 g/mol
 (C) 38.997 g/mol (D) 39.997 g/mol
39. Which of the following is not an isotope of Argon?
 (A) Argon-40 (B) Argon-36
 (C) Argon-32 (D) Argon-28
40. Oxygen element belongs to which group of the periodic table-
 (A) Triels (B) Tetrels
 (C) Pnictogens (D) Chalcogens
41. What is the boiling point of ethanol?
 (A) 77.67°C (B) 78.37°C
 (C) 79.36°C (D) 76.87°C
42. Which of the following is boltzman constant-
 (A) $1.38064852 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$
 (B) $1.36528954 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$
 (C) $1.32965513 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$
 (D) $1.30012355 \times 10^{-23} \text{ m}^2 \text{ kg s}^{-2} \text{ K}^{-1}$
43. Who gave the Law of conservation of mass?
 (A) Joseph Luis (B) Joseph Proust
 (C) Antoine Lavoisier (D) Avogadro
44. What is the molecular weight of glucose?
 (A) 186.154 u (B) 180.162 u
 (C) 195.235 u (D) 196.212 u
45. Find the number of neutrons in ${}_{29}^{56}\text{Fe}$
 (A) 26 (B) 56
 (C) 82 (D) 28
46. If a wave bends around an obstacle, then it is called _____.
 (A) diffraction (B) interference
 (C) refraction (D) enforcement
47. What is a black body?
 (A) An ideal body, which emits and absorbs radiations of all frequencies uniformly.
 (B) An ideal body, which emits and absorbs radiations of some specific frequencies uniformly.
 (C) An ideal body, which emits and absorbs radiations of all frequencies disorderly.
 (D) An ideal body, which emits and absorbs radiations of some specific frequencies disorderly.
48. What is the correct sequence of alpha beta and gamma rays in terms of distinction?
 (A) Gamma < Beta < Alpha
 (B) Beta < Gamma < Alpha
 (C) Alpha < Gamma < Beta
 (D) Alpha < Beta < Gamma
49. Isobars are the atoms with-
 (A) Same mass number but different atomic number.
 (B) Different mass number but same atomic number.
 (C) Mass number and Atomic number both are the same.
 (D) Atomic number and Mass number both are different.
50. The electrons are ejected from the metal surface as soon as the beam of light strikes the surface, then how much time lag between the striking of light beam and the ejection of electrons from the metal surface?
 (A) $\frac{1}{4}$ second
 (B) $\frac{1}{10}$ second
 (C) There is no time lag
 (D) $\frac{1}{25}$ second

51. Lysergic acid diethylamide (LSD) is a drug used a _____.
 (A) Steroid (B) Sedative
 (C) Analgesic (D) Hallucinogen
52. The nucleus of an atom consists of
 (A) Electrons and neutrons
 (B) Electrons and protons
 (C) Protons and neutrons
 (D) All of the above
53. The Potassium graphite and Calcium Graphite can be mostly used as which among the following?
 (A) Moderators (B) Semiconductors
 (C) Superconductors (D) Lubricants
54. The number of moles of solute present in 1 kg of a solvent is called its
 (A) Molality (B) Molarity
 (C) Normality (D) Formality
55. The most electronegative element among the following is.
 (A) Sodium (B) Bromine
 (C) Fluorine (D) Oxygen
56. Which method is used in general to obtain metal from its sulphide ore?
 (A) Reduction (B) Roasting
 (C) Calcination (D) Electrolysis...
57. The metal used to recover copper from a solution of copper sulphate is _____.
 (A) Na (B) Ag
 (C) Hg (D) Fe
58. The number of d-electrons in Fe^{2+} ($Z = 26$) is not equal to that of.
 (A) p-electrons in Ne
 (B) s-electrons in Mg
 (C) d-electrons in Fe
 (D) p-electrons in Cl
59. The metallurgical process in which a metal is obtained in a fused state is called.
 (A) Smelting (B) Roasting
 (C) Calcinations (D) Froth floatation
60. The molecules of which gas have highest speed?
 (A) H_2 at -73°C (B) CH_4 at 300 K
 (C) N_2 at $1,027^\circ\text{C}$ (D) O_2 at 0°C
61. The oldest rocks in the earth's crust that spewed out in volcanic eruptions during the earth early life and solidified are called _____.
 (A) Granite (B) Basalt
 (C) Igneous rocks (D) Sedimentary
62. The most commonly used bleaching agent is
 (A) Alcohol (B) Carbon dioxide
 (C) Chlorine (D) Sodium chlorine
63. The monomer of polythene is
 (A) Vinyl chloride (B) Ethylene
 (C) Ethyl alcohol (D) None of these
64. The luster of a metal is due to _____.
 (A) Its High density
 (B) Its High polishing
 (C) Its chemical inertness
 (D) Presence of free electrons
65. The number of water molecules present in a drop of water (volume 0.0018 ml) at room temperature is _____.
 (A) 1.568×10^3 (B) 6.023×10^{19}
 (C) 4.84×10^{17} (D) 6.023×10^{23}
66. The most malleable metal is
 (A) Platinum (B) Silver
 (C) Iron (D) Gold
67. The oil used in the froth floatation process is
 (A) Coconut oil (B) Olive oil
 (C) Kerosene oil (D) Pine oil
68. The mass of one Avogadro number of helium atom is
 (A) 1.00 gram
 (B) 4.00 gram
 (C) 8.00 gram
 (D) $4 \times 6.02 \times 10^{23} / \text{gram}$
69. The material which can be deformed permanently by heat and pressure is called a
 (A) Thermoplastic
 (B) Thermoset
 (C) Chemical compound
 (D) Polymer
70. The gas used in the manufacture of vanaspati from vegetable oil is
 (A) Hydrogen (B) Oxygen
 (C) Nitrogen (D) Carbon dioxide
71. The ionic radii of N^{3-} , O^{2-} , F^- and Na^+ follows the order
 (A) $\text{N}^{3-} > \text{O}^{2-} > \text{F}^- > \text{Na}^+$
 (B) $\text{N}^{3-} > \text{Na}^+ > \text{O}^{2-} > \text{F}^-$
 (C) $\text{Na}^+ > \text{O}^{2-} > \text{N}^{3-} > \text{F}^-$
 (D) $\text{O}^{2-} > \text{F}^- > \text{Na}^+ > \text{N}^{3-}$
72. Which of these is a hydronium ion?
 (A) H^+ (B) HO^-
 (C) H_2^+ (D) H_3O^+

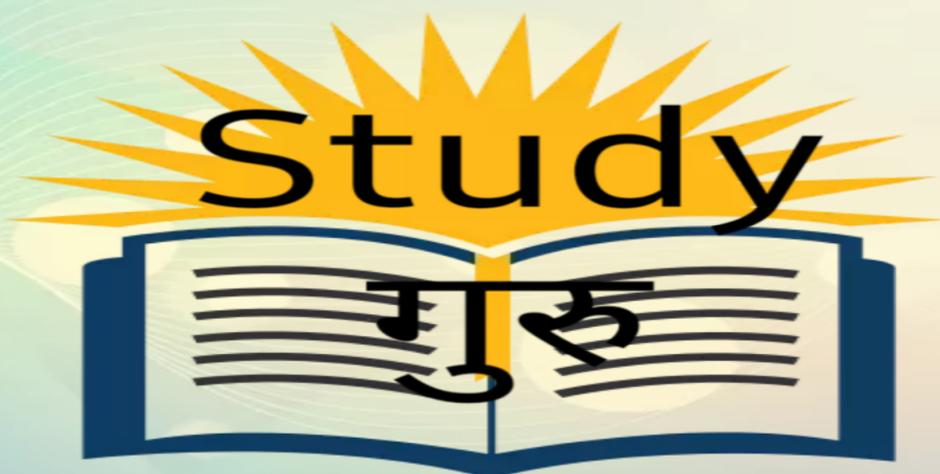
73. The most electropositive elements among the following is
 (A) Na (B) Ca
 (C) K (D) Cs
74. The gas used for artificial ripening of green fruit is_____.
 (A) Ethylene (B) Ethane
 (C) Carbon dioxide (D) Acetylene
75. Zone refining is used for the purification of
 (A) Au (B) Ge
 (C) Ag (D) Cu
76. Lysergic acid diethylamide (LSD) is a drug used as a :
 (A) Steroid (B) Sedative
 (C) Analgesic (D) Hallucinogen
77. "Oil of mirbane" is the most common name of which among the following?
 (A) Phenol (B) Toluene
 (C) Phenolphthalein (D) Nitrobenzene
78. The number of moles of solute present in 1 kg of a solvent is called its _____.
 (A) Molality (B) Molarity
 (C) Normality (D) Formality
79. Who is the most electro negative element?
 (A) Sodium (B) Bromine
 (C) Fluorine (D) Oxygen
80. The metal used to recover copper from a solution of copper sulphate is _____.
 (A) Na (B) Ag
 (C) Hg (D) Fe
81. The metallurgical process in which a metal is obtained in a fused state is called_____.
 (A) Smelting (B) Roasting
 (C) Calcinations (D) Froth floatation
82. The heat required to raise the temperature of body by 1 K is called_____.
 (A) Specific heat
 (B) Thermal capacity
 (C) Latent heat
 (D) None of the above
83. The octane number of zero is assigned to
 (A) 2-methyl octane (B) N-heptane
 (C) Iso-octane (D) 3-methyl octane
84. The metal that is used as a catalyst in the hydrogenation of oils is_____.
 (A) Ni (B) Pb
 (C) Cu (D) Pt
85. The ore which is found in abundance in India is_____.
 (A) Monazite (B) Fluorspar
 (C) Bauxite (D) Magnetite
86. The inherited traits of an organism are controlled by
 (A) RNA molecules (B) Nucleotides
 (C) DNA molecules (D) Enzymes
87. Name of the monomer of polythene is _____.
 (A) Vinyl chloride (B) Ethylene
 (C) Ethyl alcohol (D) None of these
88. The luster of a metal is due to_____.
 (A) High density
 (B) High polishing
 (C) Chemical inertness
 (D) Presence of free electrons
89. The graphite rods in the nuclear reactor
 (A) React with uranium to release energy
 (B) Produce neutrons
 (C) Undergo combustion which triggers the nuclear fission
 (D) Convert fast moving neutrons into thermal neutrons
90. The hydronium ion is_____.
 (A) H⁺ (B) HO⁻
 (C) H₂⁺ (D) H₃O⁺
91. The gas used for artificial ripening of green fruit is _____.
 (A) Ethylene (B) Ethane
 (C) Carbon dioxide (D) Acetylene
92. The mineral containing both magnesium and calcium is_____.
 (A) Magnetite (B) Calcite
 (C) Carnallite (D) Dolomite
93. The most extensive, commercially useful source of thorium as monazite sand occurs in India at Which place?
 (A) Orissa coast
 (B) Travancore coast
 (C) West Bengal coast
 (D) Gujarat coast
94. The number of electrons presents in H⁺ is _____.
 (A) Zero (B) One
 (C) Two (D) Three
95. The human body is made up of several chemical elements; the element present in the highest proportion (65%) in the Human body

- (A) Carbon (B) Hydrogen
(C) Oxygen (D) Nitrogen
96. The names of the scientists, Newlands, Mendeleev, and Meyer are associated with the development of _____.
(A) Atomic structure
(B) Metallurgy
(C) Periodic table of contents
(D) Discovery of elements
97. Which among the following is known as Quick Lime?
(A) CaO (B) CaCO₂
(C) Ca(OH)₂ (D) CaCl₂
98. Which among the following substances is most suitable for making Compact Discs?
(A) PVC (B) Polyethylene
(C) Polyamides (D) Polycarbonates
99. What is the common name of analgesic and antipyretic drug acetylsalicylic acid?
(A) Paracetamol (B) Aspirin
(C) Wintergreen (D) Trazodone
100. Rayon is an example of _____.
(A) A synthetic fiber
(B) A natural fiber
(C) Semi-synthetic fiber
(D) Synthetic Detergent
101. Where the Bio-chemical compounds are used?
(A) Skin Treatments
(B) Food preservatives
(C) Cooking Oils
(D) All of the above
102. What is the composition of soap?
(A) Sodium salt with fatty acids.
(B) Potassium salt with fatty acids
(C) Both a & b
(D) Sodium and Potassium salt mixed with chemicals
103. Which is not used as an alkali?
(A) Sodium hydroxide
(B) Potassium hydroxide
(C) Carbon hydroxide
(D) Nitrogen hydroxide
104. Which acid is present in lemon?
(A) malic acid (B) citric acid
(C) lactic acid (D) tartaric acid
105. How will you define the process of Vulcanization?
(A) Sample of butane mixed with sulphur and litharge
(B) Sample of propane mixed with sulphur and litharge
(C) Sample of plastic formed carbon mixed with sulphur and litharge
(D) Sample of rubber mixed with sulphur and litharge
106. A synthetic rubber is having _____ weight.
(A) Higher resistance (B) Lower density
(C) Higher molecular (D) Higher atomic
107. Which of the following is not a type of elements?
(A) Metals (B) Non-Metals
(C) Metalloids (D) Gases
108. Which of the following is the heaviest metal?
(A) osmium (B) mercury
(C) iron (D) nickel
109. Which of the following is not greenhouse gas?
(A) CO (B) O₃
(C) CH₄ (D) H₂O Vapour
110. Which of the following treatment is used for removal of biological impurities?
(A) Sedimentation (B) Boiling
(C) Sterilization (D) distillation
111. Which of the following indicator used in determination of hardness?
(A) EBT (B) Phenolphthalein
(C) Methyl orange (D) Thymol blues
112. Who is known as the father of Modern Chemistry?
(A) Kolvey (B) Wholer
(C) Leviatiae (D) Pasteur
113. The inventor of the atomic theory is:
(A) Rutherford (B) Madam Curie
(C) John Dalton (D) Albert Einstein
114. Nucleon is the name which is employed for:
(A) Electron and proton
(B) proton and neutron
(C) Electron and neutron
(D) None of these
115. The mass number of an atom is:
(A) Number of nucleons in the nucleus
(B) Number of protons in the nucleus
(C) Number of neutrons in the nucleus
(D) None of these
116. The number of neutrons and protons in the nucleus of ⁸⁸Ra²²⁶ are:
(A) 138 and 88
(B) 88 and 138
(C) 226 and 88
(D) 88 and 226
117. Who was the inventor of radioactivity?
(A) Madam Curie (B) Irine Curie
(C) Henery Bacqurel (D) Rutherford
118. Which of the following is negatively charged?
(A) Alpha-rays (B) Beta-rays
(C) Gama-rays (D) X-ray

119. Which of the following is not a radioactive element?
 (A) Astatine (B) Francium
 (C) Titanium (D) Zirconium
120. The number of isotopes in the hydrogen:
 (A) 2 (B) 3
 (C) 4 (D) 5
121. Which of the following is the unit of radioactivity?
 (A) Curie (B) Becquerel
 (C) Rutherford (D) All of these
122. What is the process called when a compound combines with oxygen gas to form water, heat and carbon dioxide?
 (A) Electrolysis (B) Combustion
 (C) Diffusion (D) Calorimetry
123. When a gas moves through an opening into a chamber that contains no pressure, it is called:
 (A) Effusion (B) Diffusion
 (C) Combustion (D) Osmosis
124. When a substance loses electrons, it is called:
 (A) Corrosion (B) Rust
 (C) Oxidation (D) Osmosis
125. The organic compounds having the same molecular formula but different structures are called:
 (A) Atoms
 (B) Isomers
 (C) Homologous series
 (D) Haloalkanes
126. Which of the following metal (shown by its symbol) is generally used for making filaments of bulb?
 (A) Cu (B) Pb
 (C) W (D) Ag
127. Which amongst the following is not a Cation?
 (A) Potassium ion (B) Sodium ion
 (C) Hydrogen ion (D) Sulphate ion
128. Process of gaining electrons is known as _____.
 (A) Reduction
 (B) Oxidation
 (C) Both oxidation and reduction
 (D) None of the above
129. Who gives the modern classification of periodic table?
 (A) Rutherford
 (B) D.I. Mendeleev
 (C) Moseley
 (D) Bronsted and Lowry
130. Which element in the periodic table is greater atomic number than sulphur?
 (A) Chlorine (B) Aluminum
 (C) Carbon (D) Magnesium
131. Which of the following elements are commonly found in most fertilizers?
 (A) Sodium, Potassium, Phosphorus
 (B) Sodium, Potassium, Calcium
 (C) Nitrogen, Potassium, Phosphorus
 (D) Nitrogen, Potassium, Calcium
132. Which of the following is used as moderator in atomic reactor?
 (A) Phosphorous (B) Thorium
 (C) Graphite (D) Magnesium
133. 'Oil of vitriol' is the common name of which of the following?
 (A) Picric Acid (B) Citric Acid
 (C) Acetic Acid (D) Sulphuric Acid
134. Which among the following acid is also known as 'Muriatic Acid'?
 (A) Sulfurous Acid (B) Oxalic Acid
 (C) Formic Acid (D) Hydrochloric acid
135. What is a Vermicompost?
 (A) Organic fertilizer (B) Inorganic fertilizer
 (C) Toxic Substance (D) Type of soil
136. Which one of the following gas is not found in the atmosphere?
 (A) Argon (B) Krypton
 (C) Radon (D) Xenon
137. Which one of the following is not a mixture?
 (A) Graphite (B) Glass
 (C) Brass (D) Steel
138. Select the correct match using the code given below:
- | Column-I | Column-II |
|---------------------|------------------------|
| A. Fertilizer | 1. King of Chemicals |
| B. Sulphuric acid | 2. Basic |
| C. Lime water | 3. Magnesium hydroxide |
| D. Milk of magnesia | 4. Potassium nitrate |
| A B C D | |
| (A) 1 2 3 4 | (B) 4 1 2 3 |
| (C) 3 2 1 4 | (D) 2 1 3 4 |
139. The metal compound commonly found in Sindhoor or kumkum is based on-
 (A) Tin (B) Lead
 (C) Copper (D) Zinc
140. Which one of the following metals is less reactive than hydrogen?
 (A) Barium (B) Copper
 (C) Lead (D) Magnesium
141. The bonds formed by the partnership between the electrons couples between the atoms are called -
 (A) Ionic bond (B) Covalent bond
 (C) Co-ordinate bond (D) None of the above
142. The substance made with Iron is protected from rusting by galvanization. The protecting metal used is
 (A) Silver (B) Copper
 (C) Zinc (D) Nickel

143. Which one of the following elements exhibits the greatest tendency to lose electrons?
(A) Fluorine (B) Lithium
(C) Oxygen (D) Zinc
144. Vinegar is used as a condiment, and in the pickling of vegetables and other foods. What is the constituent of vinegar?
(A) Methanoic acid (B) Tartaric acid
(C) Ethanoic acid (D) Hexanoic acid
145. Which one among the following polymers is used for making bullet-proof material?
(A) Polyethylene (B) Polyvinyl siloxane
(C) Poly toluene (D) Polyamide
146. Which one among the following substances evolved heat when dissolved in water?
(A) Lactic acid (B) Fructose
(C) Quick lime (D) Salt peter
147. At constant temperature, the product of pressure and volume of a given amount of a gas is constant. This is
(A) Gay-Lussac law (B) Charles's law
(C) Boyle's law (D) Pressure law
148. An ideal fuel should have-
(A) High calorific value
(B) Low ignition temperature
(C) Regulated and controlled
(D) All of the above
149. Which of the following cannot be beaten into Sheets?
(A) Gold (B) Silver
(C) Potassium (D) Aluminum
150. Which of the following is a radioactive metal?
(A) Iodine (B) Radium
(C) Chromium (D) Lithium

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Answer Key

1.(C)	2.(B)	3.(A)	4.(A)	5.(A)	6.(A)	7.(A)	8.(D)	9.(C)	10.(A)
11.(B)	12.(D)	13.(B)	14.(A)	15.(B)	16.(C)	17.(A)	18.(D)	19.(B)	20.(D)
21.(A)	22.(C)	23.(B)	24.(A)	25.(A)	26.(C)	27.(C)	28.(B)	29.(B)	30.(C)
31.(D)	32.(A)	33.(C)	34.(A)	35.(B)	36.(C)	37.(B)	38.(D)	39.(D)	40.(D)
41.(B)	42.(A)	43.(C)	44.(B)	45.(D)	46.(A)	47.(A)	48.(D)	49.(A)	50.(C)
51.(D)	52.(C)	53.(C)	54.(A)	55.(C)	56.(B)	57.(D)	58.(D)	59.(A)	60.(A)
61.(C)	62.(C)	63.(B)	64.(D)	65.(B)	66.(D)	67.(D)	68.(D)	69.(B)	70.(A)
71.(A)	72.(D)	73.(D)	74.(A)	75.(B)	76.(D)	77.(D)	78.(A)	79.(C)	80.(D)
81.(A)	82.(B)	83.(B)	84.(A)	85.(A)	86.(C)	87.(B)	88.(D)	89.(D)	90.(D)
91.(A)	92.(D)	93.(B)	94.(A)	95.(C)	96.(A)	97.(A)	98.(D)	99.(B)	100.(C)
101.(B)	102.(C)	103.(B)	104.(B)	105.(D)	106.(C)	107.(D)	108.(A)	109.(A)	110.(C)
111.(A)	112.(C)	113.(C)	114.(B)	115.(A)	116.(A)	117.(C)	118.(B)	119.(C)	120.(B)
121.(B)	122.(B)	123.(A)	124.(C)	125.(B)	126.(C)	127.(D)	128.(A)	129.(B)	130.(A)
131.(C)	132.(C)	133.(D)	134.(D)	135.(A)	136.(C)	137.(A)	138.(B)	139.(B)	140.(B)
141.(B)	142.(C)	143.(B)	144.(C)	145.(A)	146.(C)	147.(C)	148.(D)	149.(B)	150.(D)